The City School

**North Nazimabad Boys Campus**

**Second Monthly Test Session 2019 – 20**

**Class - 10**

**Time: 35 Minutes Chemistry Marks 30**

**Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Sec: \_\_\_\_\_\_ Date: \_\_\_\_\_\_\_\_\_\_\_\_\_\_**

Q.no1) calculate the relative molecular mass following substances

* H2O
* Na2CO3
* Al2O3
* C4H10O

  **[4 marks]**

 Q.no2

Ethane is hydrocarbon which is combustible in air to produce CO2 and water

 **C2H6 + 3.5 O2 → 2CO2 + 3H2O**

What volume of Oxygen gas is needed to completely burn 20 cm3 of ethane C2H6

Volume of O2 …………………. dm3

 **[Total marks 3]**

**Q.N03**) Calculate the mass of Cu(II)SO4 formed when 5.0 g of CuO is reacted with excess sulfuric acid H2SO4

 **CuO(S) + H2SO4 (aq) →CuSO4 (aq) + H2O (l)**

 **[Total marks 2]**

**z**

Q.NO 4)

9.30 g of sodium hydrogen carbonate was heated to form 1.22 dm3 of CO2 gas at r.t.p. calculate the percentage purity of sodium hydrogen carbonate.

 **NaHCO3(s) → Na2CO3(s) +CO2 (g) + H2O**

 **[total marks 4]**

**Q.NO 5) THIS QUESTION IS FROM MOLE ATP**

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