Teacher Name: Uzma Amer Class: 9 Subject: Chemistry Date:17-11-16

Q.1. Choose the best answers:

1. Which electron arrangement is that of a metallic element?
2. 2,1 b) 2,7 c) 2,5 d) 2,4

II .Which of the following contains the same number of electrons as an atom of neon?

a) Cl- b) Li+ c) C O-2 d) Na

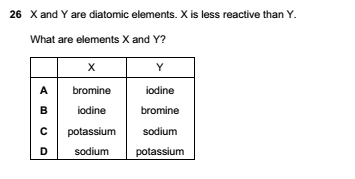
III. Use the Periodic Table to decide which element has all four of the properties shown.

\* high melting point \* good electrical conductivity

a) Caesium, Cs b) Chlorine Cl c) Iodine, I d) Strontium, Sr

1. X and Y are diatomic elements. X is less reactive than Y.

What are elements X and Y?



Element X is solid at room temperature.

It needs one electron per atom to gain the electronic structure of a noble gas.

It is the least reactive element in its group.

What is the element X?

a) At b) Cs c) F d) Li

1. The letters X, Y and Z represent different atoms.

|  |  |  |
| --- | --- | --- |
| atoms | Proton number | Nucleon number |
| X | 19 | 40 |
| Y | 19 | 39 |
| Z | 20 | 40 |

What can be deduced from the proton numbers and nucleon numbers of X, Y and Z?

1. X and Y are the same element.
2. X and Z are the same element.
3. X has more protons than Y.
4. Z has more neutrons than Y

Q2. Some atoms denoted by the letter A to E, have proton numbers which are

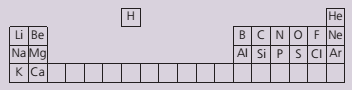
A = 3, B = 10, C = 9, D = 17, E = 11.

a) Which of these have complete outermost shell?

b) Which of these are in Group I of Periodic table?

c) Which of these are in Group VII of the Periodic table

Q3. This extract from the Periodic Table shows the symbols for the first 20 elements.



Look at the row from lithium (Li) to neon (Ne).

1. What is this row of the Periodic Table called?

b ) Which element in it is the least reactive? Why?

Look at the column of elements from lithium (Li) to potassium (K).

c) What is this column of the table called?

d) Of the three elements shown in this column, which one is the most reactive?

Q.4.Give scientific reasons:

1. Mg ion is smaller than Mg atom.
2. Non metals make Anion

Q.5. Define isotopes Draw atomic structures of isotopes of Hydrogen.