**Worksheet#1**

**Chemistry class 9**

Teacher Name: Uzma Amer Class: 9 Subject: chemistry Date: 14th April’17

Q.1 (a) Define the following terms:

1. Kinetic Energy
2. Relative molecular mass
3. Electron negativity
4. Ionizing Energy

(b).calculate the Relative molecular masses of the given compounds:

1. A compound XCl having molecular mass (Mr) of 58.5.Find out Mr of XCO3.
2. A compound X2O having molecular mass (Mr) of 62 Find out Mr of X2SO4
3. A compound XCl2 having molecular mass (Mr) of 111.Find out Mr of X(NO3)2

Q.3.Look at following diagrams and answers the questions given below:

 

1. Which atom has a mass of 5?
2. Which atoms are isotopes of each other?
3. What is the charge on above atoms?
4. Explain your answer of part c.

**Worksheet#2**

**Chemistry class 9**

Q.1. Choose the best answers:

1. Metals are solids except
2. helium
3. selenium
4. mercury
5. Potassium
6. Potassium Oxide is formed through
7. two potassium ions and one oxygen ion
8. one potassium ion and one oxygen ion
9. two potassium ions and two oxygen ion
10. one potassium ion and two oxygen ion
11. In forming an ionic bond in Potassium Fluoride (KF), Potassium ion (K+)
12. gets a single positive charge
13. gets a double positive charge
14. gets a single negative charge
15. gets a double negative charge
16. When an acid (H+) is added to alkali (OH-), product is
17. hydroxides
18. water
19. salts
20. hydrogen gas
21. Due to mobile valence electrons,
22. metals are rigid
23. metals are not rigid
24. metals are poor electrolytes
25. metals are poor insulators
26. K2O + H2O →
27. K(OH)3
28. KOH
29. KOH.H2O
30. KO + H2 + O2
31. Excess acidity caused by acid rain can be neutralized by adding
32. more fertilizers
33. by removing acidified soil
34. by adding P2O5
35. by adding limeswer D
36. More corrosive acid is
37. H2CO3
38. H2SO3
39. HNO3
40. C6H8O7

Q.2. Differentiate between:

1. Base and Alkali
2. Basic and Acidic oxides
3. Ionic and Covalent bonding

**Worksheet#3**

**Chemistry class 9**

Q.1. choose the best answers:

1. Brass is an example of
2. An alloy
3. a compound
4. a mixture
5. an elementAnswer A
6. A particle may be
7. an atom
8. a molecule
9. an isotope
10. all of theseAnswer D

III. Other than O and C, main elements of a brick wall includes

1. Silicon (Si)
2. Calcium (Ca)
3. Aluminum (Al)
4. Both A and BAnswer D

IV. Symbol of Magnesium is

1. Mg
2. Mn
3. Hg
4. M

Answer A

V.. Naturally occurring elements are

1. 82
2. 92
3. 99
4. above 100 by nowAnswer B

VI. Electronic thermometers

1. are mercury in glass thermometer
2. alcohol in glass thermometer
3. are used to reduce risk of mercury poisoning if broken
4. are greatly replaced by alcohol in glass thermometerAnswer C

**VII.** Unit for volumes does not include

1. ccm
2. cdm
3. liters
4. tonesAnswer D

VIII. 100 cm3, 250 cm3 and 1 dm3 can be measured using

1. measuring cylinder
2. beaker
3. volumetric flask
4. burette

Q.2. Cite four differences between compounds and mixtures:

**Worksheet#4**

**Chemistry class 9**

I. Chemical formula of Zinc Hydroxide is

1. Zn(OH)2
2. ZNO2
3. Zn2(OH)3
4. ZnOH
5. Valency of Carbon (C) is
6. 1
7. 2
8. 3
9. 4
10. Chemical formula of silver bromide is
11. Ag2Br3
12. AgBr
13. AgBrO
14. AgBr3.H2O
15. Crude oil can be fractionally distillated to produce
16. diesel
17. petrol
18. paraffin
19. all of these
20. Mixtures have
21. a range of boiling points
22. a range of melting points
23. a lack of exact concentrations
24. all of these

Q.2. Look at the given diagram



1. Identify at least five equipments and write down their names:
2. Describe the uses of those equipments: