

*The City School*  
Curriculum Distribution Chart

Subject: Chemistry 2019-2020  
Scheme of work/Term wise syllabus breakup  
Class Level: 9

Term 1				
STRAND	UNIT	TOPIC	OBJECTIVE	WEEK
	2.The particulate nature of matter	2.1 Kinetic particle theory	(a) describe the solid, liquid and gaseous states of matter and explain their interconversion in terms of the kinetic particle theory and of the energy changes involved (b) describe and explain evidence for the movement of particles in liquids and gases (the treatment of Brownian motion is not required) (c) explain everyday effects of diffusion in terms of particles, e.g. the spread of perfumes and cooking aromas; tea and coffee grains in water (d) state qualitatively the effect of molecular mass on the rate of diffusion and explain the dependence of rate of diffusion on temperature (e) state qualitatively and explain the effects of temperature and pressure on the volumes of gases	1

		2.3 Structure and Properties of matter and material	<p>(a) describe the differences between elements, compounds and mixtures</p> <p>(b) compare the structure of simple molecular substances, e.g. methane, iodine, with those of giant covalent substances, e.g. sand (silicon dioxide), diamond, graphite in order to deduce their properties</p> <p>(c) compare the bonding and structures of diamond and graphite in order to deduce properties such as electrical conductivity, lubricating or cutting action (candidates will not be required to draw the structures)</p> <p>(d) deduce the physical and chemical properties of substances from their structures and bonding and <i>vice versa</i></p>	1
	1.-Experimental Chemistry	1.1- Experimental design	<p>(a) name appropriate apparatus for the measurement of time, temperature, mass and volume, including burettes, pipettes, measuring cylinders and gas syringes</p> <p>(b) suggest suitable apparatus, given relevant information, for a variety of simple experiments, including collection of gases and measurement of rates of reaction</p>	2
		1.2- Methods of purification and analysis	<p>(a) describe methods of purification by the use of a suitable solvent, filtration and crystallisation, distillation and fractional distillation, with particular references to the fractional distillation of crude oil, liquid air and mixtures of liquids</p> <p>(b) suggest suitable methods of purification, given information about the substances involved</p> <p>(c) describe paper chromatography and interpret chromatograms including comparison with 'known' samples and the use of <math>R_f</math> values</p>	2

			<p>(d) explain the need to use locating agents in the chromatography of colourless compounds</p> <p>(e) deduce from given melting point and boiling point information, the identities of substances and their purity</p> <p>(f) explain that the measurement of purity in substances used in everyday life, e.g. foodstuffs and drugs, is important</p>	
	2.-The particulate nature of matter	2.2-Atomic Structure	<p>(a) state the relative charges and approximate relative masses of a proton, a neutron and an electron</p> <p>(b) describe, with the aid of diagrams, the structure of an atom as containing protons and neutrons (nucleons) in the nucleus and electrons arranged in shells (energy levels) (no knowledge of s, p, d and f classification will be expected; a copy of the Periodic Table will be available in Papers 1 and 2)</p> <p>(c) define proton number and nucleon number</p> <p>(d) interpret and use symbols such as <math>{}_6\text{C}^{3-}</math></p> <p>(e) define the term <i>isotopes</i></p> <p>(f) deduce the numbers of protons, neutrons and electrons in atoms and ions from proton and nucleon numbers</p>	2
	8.-Periodic Table	8.1-Periodic trends	<p>(a) describe the Periodic Table as an arrangement of the elements in the order of increasing proton (atomic) number</p> <p>(b) describe how the position of an element in the Periodic Table is related to proton number and electronic configuration</p> <p>(c) describe the relationship between group number and the ionic charge of an element</p> <p>(d) explain the similarities between the elements in the same group of the Periodic Table in terms of their electronic configuration</p>	2

			<p>(e) describe the change from metallic to non-metallic character from left to right across a period of the Periodic Table.</p> <p>(f) describe the relationship between group number, number of valence electrons and metallic/non-metallic character</p> <p>(g) predict the properties of elements in Group I, VII and the transition elements using the Periodic Table</p>	
		<b>8.2-Group properties</b>	<p>(a) describe lithium, sodium and potassium in Group I (the alkali metals) as a collection of relatively soft, low-density metals showing a trend in melting point and in their reaction with water</p> <p>(b) describe chlorine, bromine and iodine in Group VII (the halogens) as a collection of diatomic non-metals showing a trend in colour, state and their displacement reactions with solutions of other halide ions</p> <p>(c) describe the elements in Group VIII (the noble gases, also known as Group 0) as a collection of monatomic elements that are chemically unreactive and hence important in providing an inert atmosphere, e.g. argon and neon in light bulbs; helium in balloons; argon in the manufacture of steel</p> <p>(d) describe the lack of reactivity of the noble gases in terms of their electronic configuration</p>	2
		<b>8.3-Transition elements</b>	<p>a) describe the central block of elements (transition metals) as metals having high melting points, high density, variable oxidation state and forming coloured compounds</p> <p>b ) state the use of these elements and/or their compounds as catalysts, e.g. iron in the Haber process; vanadium(V) oxide in the contact process; nickel in the hydrogenation of alkenes, and how catalysts are used in industry to lower energy demands and hence are economically advantageous and help conserve energy sources</p>	1

<b>Term 2</b>				
	2.-The particulate nature of matter	2.4-Ionic bonding	<p>(a) describe the formation of ions by electron loss/gain in order to obtain the electronic configuration of an inert gas</p> <p>(b) describe the formation of ionic bonds between metals and non-metals, e.g. NaCl; <math>MgCl_2</math></p> <p>© state that ionic materials contain a giant lattice in which the ions are held by electrostatic attraction, e.g. NaCl (candidates will not be required to draw diagrams of ionic lattices)</p> <p>(d) relate the physical properties (including electrical property) of ionic compounds to their lattice structure</p>	2
		2.5-Covalent bonding	<p>(a) describe the formation of a covalent bond by the sharing of a pair of electrons in order to gain the electronic configuration of an inert gas</p> <p>(b) describe, using dot-and-cross diagrams, the formation of covalent bonds between non-metallic elements, e.g. <math>H_2</math>; <math>Cl_2</math>; <math>O_2</math>; HCl; <math>N_2</math>; <math>H_2O</math>; <math>CH_4</math>; <math>C_2H_4</math>; <math>CO_2</math></p> <p>© deduce the arrangement of electrons in other covalent molecules</p> <p>(d) relate the physical properties (including electrical properties) of covalent compounds to their structure and bonding</p>	2
		2.6-Metallic Bonding	<p>(a) describe metallic bonding as the electrostatic attraction between positive ions in a lattice and a 'sea of electrons'</p> <p>(b) relate the malleability of metals to their structure and the electrical conductivity of metals to the mobility of the electrons in the structure</p>	1
	3-Formulae, stoichiometry and the mole concept	Formula and Equation Making	<p>a) state the symbols of the elements and formulae of the compounds mentioned in the syllabus</p>	3

			<p>(b) deduce the formulae of simple compounds from the relative numbers of atoms present and <i>vice versa</i></p> <p>(c) deduce the formulae of ionic compounds from the charges on the ions present and <i>vice versa</i></p> <p>(d) interpret and construct chemical equations, with state symbols, including ionic equations</p>	
	7.-The chemistry and uses of acids, bases and salts	7.1- The characteristic properties of acids and bases	<p>(a) describe the meanings of the terms acid and alkali in terms of the ions they contain or produce in aqueous solution and their effects on Universal Indicator paper</p> <p>(b) describe how to test hydrogen ion concentration and hence relative acidity using Universal Indicator paper and the pH scale</p> <p>(c) describe the characteristic properties of acids as in reactions with metals, bases and carbonates</p> <p>(d) describe qualitatively the difference between strong and weak acids in terms of the extent of dissociation</p> <p>(e) describe neutralisation as a reaction between hydrogen ions and hydroxide ions to produce water, <math>\text{H}^+ + \text{OH}^- \rightarrow \text{H}_2\text{O}</math></p> <p>(f) describe the importance of controlling the pH in soils and how excess acidity can be treated using calcium hydroxide</p> <p>(g) describe the characteristic properties of bases in reactions with acids and with ammonium salts</p> <p>(h) classify oxides as acidic, basic or amphoteric, based on metallic/non-metallic character</p>	2
		7.2-Preparation of salts	<p>(a) describe the techniques used in the preparation, separation and purification of salts as examples of some of the techniques specified in Section 1.2(a)</p>	3

			<p>(methods for preparation should include precipitation and titration together with reactions of acids with metals, insoluble bases and insoluble carbonates)</p> <p>(b) describe the general rules of solubility for common salts to include nitrates, chlorides (including silver and lead), sulfates (including barium, calcium and lead), carbonates, hydroxides, Group I cations and ammonium salts</p> <p>(c) suggest a method of preparing a given salt from suitable starting materials, given appropriate information</p> <p>(d) describe the meanings of the terms hydrated, anhydrous and water of crystallisation</p>	
	1.-Experimental Chemistry	1.3 Identification of ions and gases	<p>(a) describe the use of aqueous sodium hydroxide and aqueous ammonia to identify the following aqueous cations: aluminium, ammonium, calcium, chromium(III), copper(II), iron(II), iron(III) and zinc (formulae of complex ions are <b>not</b> required)</p> <p>(b) describe tests to identify the following anions: carbonate (by the addition of dilute acid and subsequent use of limewater); chloride (by reaction of an aqueous solution with nitric acid and aqueous silver nitrate); iodide (by reaction of an aqueous solution with nitric acid and aqueous silver nitrate); nitrate (by reduction with aluminium and aqueous sodium hydroxide to ammonia and subsequent use of litmus paper), sulfate (by reaction of an aqueous solution with nitric acid and aqueous barium nitrate), and sulfite (by warming with dilute acid)</p> <p>(c) describe tests to identify the following gases: ammonia (using damp red litmus paper); carbon dioxide (using limewater); chlorine (using damp litmus paper); hydrogen (using a burning splint); oxygen (using a glowing splint) and sulfur dioxide (using acidified potassium manganate(VII))</p>	2

			(d) describe a chemical test for water	
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