

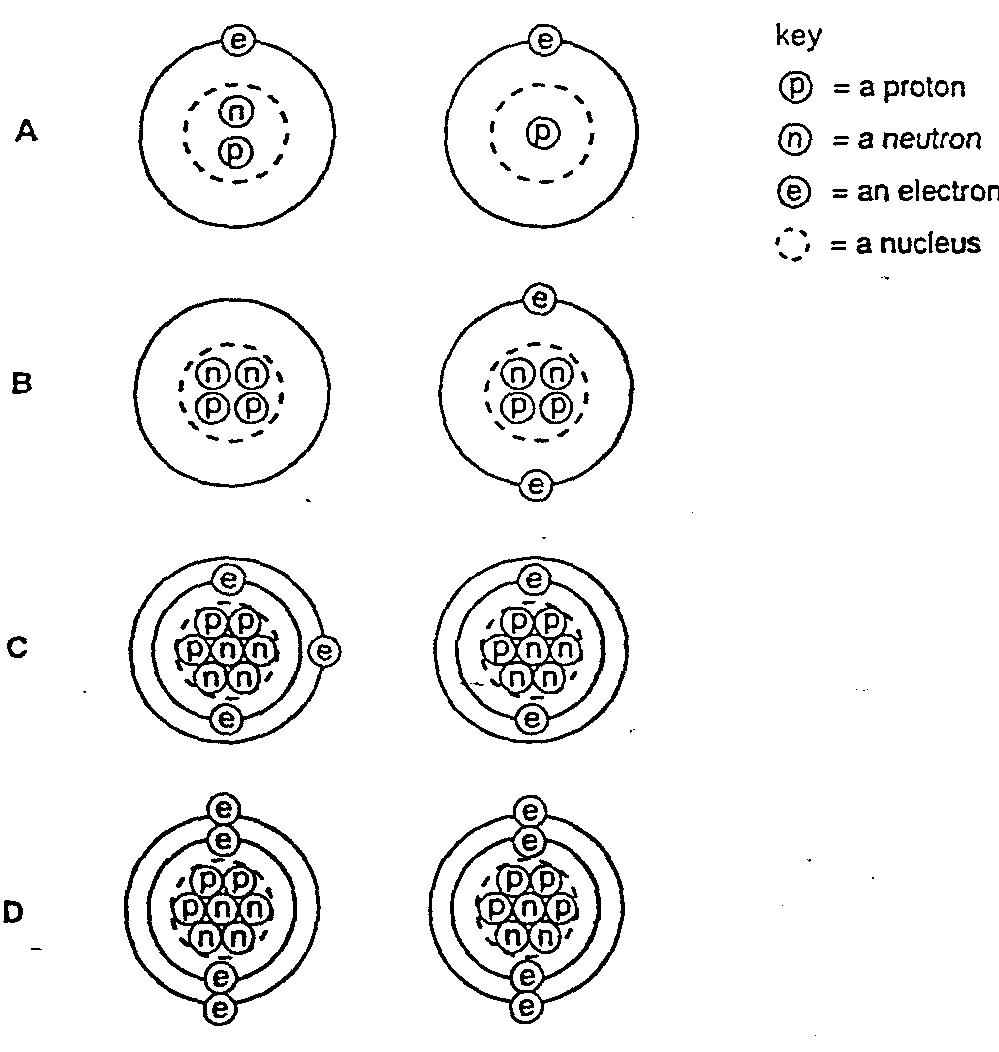
**CHEMISTRY CLASS 9**

**REINFORCEMENT WORKSHEET**

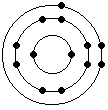
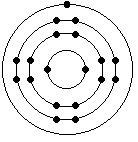
**Ms. UZMA AMIR**

**Date: 15-5-2015**

1. Which two diagrams show two different types of atom of the same element?

[](http://3.bp.blogspot.com/-06Q6WHVPlP4/UZ-BqNpal3I/AAAAAAAABWA/PVW4bHIPBTU/s1600/s986.jpg)

1. Look at given atomic structures

Wblue.gifXblue.gifY.gifZblue.gif  
Which electron arrangement corresponds to the first element on period 2 of the Periodic Table?

1. electron arrangement X
2. electron arrangement **Y**
3. electron arrangement **W**
4. electron arrangement **Z**
5. **look at the periodic table;**

  
Where in the Periodic Table are you MOST LIKELY to find an element that is a low melting solid and a good conductor of heat?

1. Group 7
2. Group 1
3. Transition series
4. Group 4
5. Which four quantities **A**, **B**, **C** and **D** are required to balance the equation? (Note: 1's are not written in the equation BUT they are needed in the 'balancing ratio thinking'.)

**A** Na2O(s) + **B** H2SO4(aq) ==> **C** Na2SO4(aq) +**D** H2O(l)

1. 1 2 1 2
2. 2 2 2 1
3. 2 1 2 2
4. 1 1 1
5. Which is **TRUE** about the compound formed on combining an Alkali Metal with the Halogen element chlorine to form the metal chloride?
6. the solution of it in water conducts electricity
7. the solid conducts electricity
8. insoluble in water
9. is an ionic compound
10. I is only
11. I and II only
12. I, II and III only
13. I an IV only
14. Which of the following is **TRUE** about the trend DOWN the Group 7 Halogens with increase in atomic number?
15. the colour of the element gets darker
16. the melting points decrease
17. the reactivity increases
18. the atoms get smaller

1. The equation for the reaction between aqueous lead(II) nitrate and aqueous potassium iodide is shown.

Pb(NO3) 2(aq) + 2KI(aq) → PbI2(s) + 2KNO3(aq)

Colourless colourless yellow colourless

1. Which method could be used to separate the products?

A chromatography

B crystallization

C distillation

D filtration

1. How does a magnesium atom form a bond with an oxygen atom?

A by giving one pair of electrons to the oxygen atom

B by sharing one pair of electrons, both electrons provided by the magnesium atom

C by sharing two pairs of electrons, both pairs provided by the oxygen atom

D by sharing two pairs of electrons, each atom donating one pair of electrons

1. What is the general relationship between lattice energy and ionic bond strength?
2. **The** greater the lattice energy, the weaker the ionic bond
3. there is no relationship
4. the greater the lattice energy, the stronger the ionic bond
5. the weaker the lattice energy, the stronger the ionic bond
6. Magnesium oxide may be used for the lining of an electric furnace for making crockery. Which properties of magnesium oxide help to explain this use?

|  |  |  |  |
| --- | --- | --- | --- |
|  | strong forces between particles | ionic bonding | electrical conductor |
| A  B  C  D | Yes  Yes  No  No | Yes  No  Yes  No | No  Yes  No  Yes |

1. The table shows the electronic structures of four atoms.

|  |  |
| --- | --- |
| Atom | electronic structure |
| W  X  Y  Z | 2,8,1  2,8,4  2,8,7  2,8,8 |

Which two atoms combine to form a covalent compound?

1. W and X
2. W and Y
3. X and
4. D X and Z
5. The following statement is about chemical bonds. Covalent bonds are formed by the …1… of electrons. Covalent substances have …2… electrical conductivity. Which words complete the statement?

|  |  |  |
| --- | --- | --- |
|  | 1 | 2 |
| A  B  C  D | Sharing  Sharing  Transfer  Transfer | High  Low  High  Low |

1. Below are two statements about metals.

I .Metals contain a lattice of negative ions in a ‘sea of electrons’.

II. The electrical conductivity of metals is related to the mobility of the electrons in the structure. Which is correct?

1. Both statements are correct and statement 1 explains statement 2.
2. Both statements are correct but statement 1 does not explain statement 2.
3. Statement 1 is correct and statement 2 is incorrect.
4. Statement 2 is correct and statement 1 is incorrect
5. Which of the following is **NOT** a trend that varies systematically in the periodic table?
6. Electron Negativity
7. symbols of elements
8. ionization energy
9. atomic radius
10. following picture are showing some methods of Purification