**BLOG WORKSHEET**

**CHEMISTRY CLASS 9**

Teacher Name: Uzma Amer Class: 9 Chemistry Date: 26th jan’18

Q.1. Choose the best answers:

1. Which electron arrangement is that of a metallic element?
2. 2,1 b) 2,7 c) 2,5 d) 2,4

II .Which of the following contains the same number of electrons as an atom of neon?

a) Cl- b) Li+ c) C O-2 d) Na

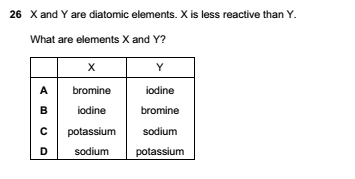
III. Use the Periodic Table to decide which element has all four of the properties shown.

\* high melting point \* good electrical conductivity

a) Caesium, Cs b) Chlorine Cl c) Iodine, I d) Strontium, Sr

1. X and Y are diatomic elements. X is less reactive than Y.

What are elements X and Y?



Element X is solid at room temperature.

It needs one electron per atom to gain the electronic structure of a noble gas.

It is the least reactive element in its group.

What is the element X?

a) At b) Cs c) F d) Li

1. The letters X, Y and Z represent different atoms.

|  |  |  |
| --- | --- | --- |
| atoms | Proton number | Nucleon number |
| X | 19 | 40 |
| Y | 19 | 39 |
| Z | 20 | 40 |

What can be deduced from the proton numbers and nucleon numbers of X, Y and Z?

1. X and Y are the same element.
2. X and Z are the same element.
3. X has more protons than Y.
4. Z has more neutrons than Y

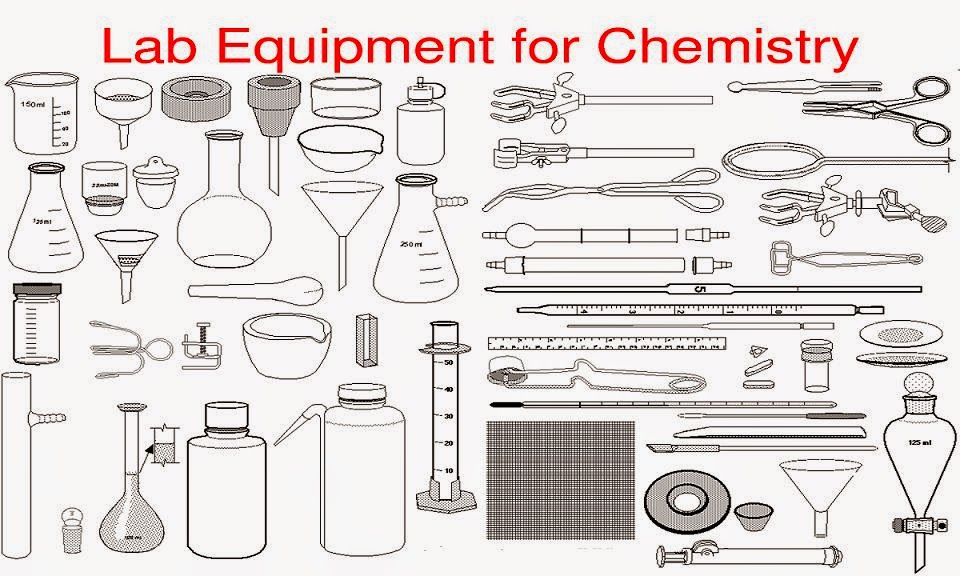
Q.2. Define the following terms:

1. Valancy
2. Oxidation State
3. Ion
4. Electrostatic force

Q.2. Draw the electronic structure of both a sodium ion and a chloride ion.[2]

Sodium ion Chloride ion

Q.3. Look at the given picture and answer the following questions:



1. Identify at least five equipments and write down their names:
2. Describe the uses of those equipments: