**Important Preparation Reinforcement worksheets**

**For Final Examination 2019**

**Chemistry 5070 paper 1**

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Q.1. Choose the best answers:

















9 Which statement explains why sodium chloride, NaCl, has a lower melting point than magnesium oxide, MgO?

A Sodium chloride is covalent but magnesium oxide is ionic.

B Sodium is more reactive than magnesium.

C The attraction between Na+ and Cl − is weaker than that between Mg2+ and O2−.

 D The melting point of sodium is lower than that of magnesium.

10. Which property decides the order of the elements in the periodic table?



11. Which feature of a metal’s structure is responsible for it conducting electricity?

 **A** It contains positive ions.

**B** It has a “sea of electrons”.

**C** Its ions are tightly packed together.

**D** Its positive ions attract electrons

12. Aqueous hydrogen peroxide undergoes catalytic decomposition as shown in the equation below.



13. A mixture of two substances is spotted on to a piece of chromatography paper.



14. Which diagram represents the structure of sand, SiO2?



15. The table gives information about three indicators.



16.



When are liquid and solid both present?

 **A** P to Q and R to S

 **B** P to Q

**C** Q to R

**D** R to S

17. Element X is a solid at room temperature.

It needs one electron per atom to gain the electronic structure of a noble gas.

 It is the least reactive element in its group.

What is the element X?

 **A** At

 **B** Cs

**C** F

 **D** Li

18. Element X and Y are in group VII of the periodic table.



19. Substance X melts at 53 o C and boils at 100 o C. It does not dissolve in water and it does not react with water.

 Which diagram shows the method most suitable for separating X from a mixture of X and water?





20. Metals have positive ions in a ‘sea of electrons’.

Which metal atom provides most electrons for the sea?

A aluminum

 B calcium

C magnesium

D sodium

21. The diagram shows the arrangement of electrons in a molecule of compound YZ2



22. Which two statements about a covalent bond are correct?

 1 It can be formed between two metal atoms.

 2 It can be formed between two non-metal atoms.

 3 It is formed by the transfer of electrons between atoms.

 4 It is formed by sharing electrons between atoms.

 A 1 and 3

 B 1 and 4

C 2 and 3

D 2 and 4

23. In which pair of substances does each have a giant molecular structure?

A diamond, iodine

B diamond, silica (sand)

 C iodine, methane

 D methane, silica (sand)

24. The diagram shows a simple laboratory apparatus for the preparation and collection of dry gas.



What is the gas?

 **A** carbon dioxide

 **B** chlorine

 **C** hydrogen

 **D** hydrogen chloride

25. The table shows the boiling points of the elements found in a sample of liquid air.



26. Naturally occurring bromine has a relative atomic mass of 80 and consists entirely of two isotopes of relative isotopic masses 79 and 81.

 What can be deduced about naturally-occurring bromine from this information only?

A Bromine isotopes have different numbers of protons.

 B Bromine contains the two isotopes in equal proportions.

 C Bromine has different oxidation states.

D Bromine is radioactive.

27. Which property shows that a liquid is pure?

A It turns anhydrous copper(II) sulphate blue.

 B It is colourless and odourless.

C It has no effect on red or blue litmus paper.

D It boils at a fixed temperature at a given pressure.

28. Which diagram represents the arrangement of particles in a gas?



Which gas diffuses at the same rate as nitrogen gas?

 A carbon dioxide

B carbon monoxide

C neon

 D sulphur dioxide

29. Which statement about diamond and graphite is correct?

A Both diamond and graphite are used as abrasives.

 B Diamond and graphite have different arrangements of carbon atoms.

C The carbon atoms in graphite have a different number of neutrons from those in diamond.

D The carbon atoms in both graphite and diamond have four covalent bonds.

30. Which diagram correctly shows the arrangement of the ions in solid sodium chloride?



 